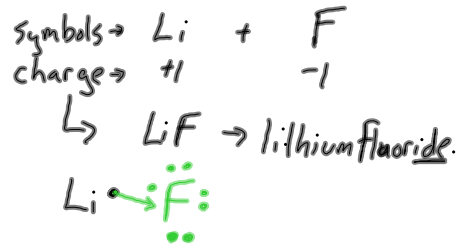


- ① Compounds.
 ↳ 2 or more elements combining together
- ② Ionic Compounds → when metal and non metal element combine (using outside valence ring and + or - charges.)
- ③ Name (how you identify a compound)
 ex Lithium + Fluorine
 lithium fluoride

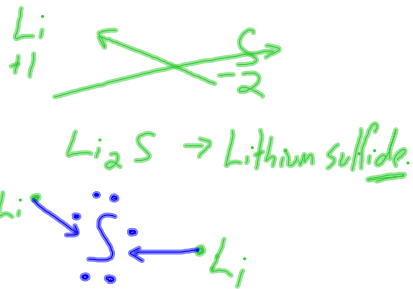
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Balanced Formulas (symbols)
 ex Lithium Fluorine.



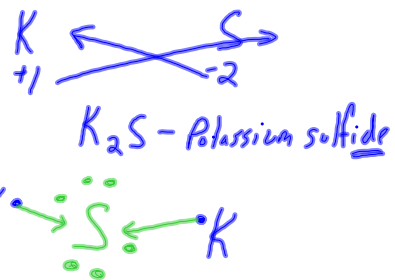
Sep 21-12:40 PM

Lithium + Sulfur



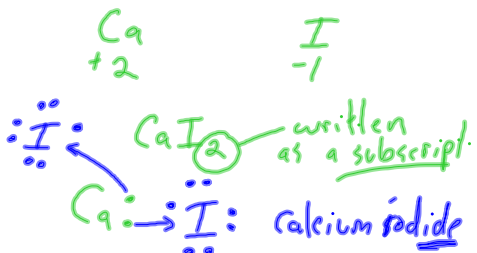
Sep 21-12:44 PM

Potassium + Sulfur



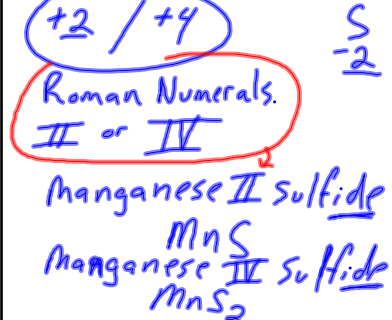
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Calcium + Iodine



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Manganese + Sulfur



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Ionic Compounds

- Metals and non-metals combine to form **ionic compounds** by transferring electrons.
- Metal atoms lose electrons to form positive ions.
- Non-metal atoms gain electrons to form negative ions.

Lewis dot diagrams
→ energy same amount of energy to gain or lose an electron

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Molecular Compounds

above slides (Sept 21)

- Molecular Compounds contain **neutral groups of atoms** called molecules.
- Covalent bonds are created when **nonmetal atoms** and their outside valence rings share electrons that hold atoms together.

Next step

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- Diatomic molecules are made from **two similar atoms** creating a covalent bond.

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Example:	Symbol	Formula	State
Hydrogen	H	H ₂	gas
Oxygen	O	O ₂	gas
Nitrogen	N	N ₂	gas
Fluorine	F	F ₂	gas
Chlorine	Cl	Cl ₂	gas
Bromine	Br	Br ₂	liquid
Iodine	I	I ₂	solid

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- Atoms of different elements can also form covalent bonds.
- Combining capacity of a nonmetal is a measure of the number of covalent bonds that it will need to form a stable molecule. This number replaces the ionic charge for writing formulas.

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